

## CHEM 162, Section 1, Spring 2012

### GENERAL CHEMISTRY II

**Instructor:** Prof. Oscar Navarro

**Contact:** office: Bilger 236  
email: oscaranf@hawaii.edu

**Classroom:** Bilger 152

**Hours:** 14:30-15:20 MWF, 3 credit hours

**Textbook:** "Chemistry: A Molecular Approach", Nivaldo J. Tro, Custom Edition  
ISBN 10: 0-558-29865-6  
Solutions Manual ISBN 10: 0-558-26010-1

**Office hours:** Students are encouraged to meet with the instructor for questions, additional information or any other related matter. Office hours are MWF 13:15-14:15, no appointment required. The instructor will make his best to accommodate alternative times if needed (request by email or in person). There will be no office hours on exam days. In addition, Bilger 337 will be staffed by Teaching Assistants Monday through Friday. Check the posted schedule for the times when TA's are available.

**Student Responsibility:** It is the student's responsibility to put forth the effort required to learn the material and to become competent with it. The best way of learning is to self-test what you don't know and correct that: this means working during the course, reviewing in a daily basis and using good study habits. Ignoring the subject until three days before the test usually guarantees a failure. If you find yourself having trouble catching up with the subject, the instructor will be happy to help you. Consult him as soon as possible. Reading the corresponding chapter and trying to understand the concepts before going to the classroom is highly encouraged and recommended.

#### Course Policies:

- 1- There will be no makeup exams. If you miss an exam and have a valid excuse (doctor's note or equivalent), the weighting of the other exams will be adjusted accordingly. Only **one** exam can be missed during the course and the final exam **cannot** be missed.
- 2- Regular attendance in lecture is highly recommended, but not mandatory. The aim of the lecture session is to guide you in your studies and to clarify, emphasize and illustrate the important (and sometimes subtle) concepts. Topics not included in the text will be covered in class and will appear in the tests. You are responsible for all information relayed in class whether you attend or not.
- 3- Academic dishonesty will not be tolerated. Cheating in the form of copying, plagiarism, altering information, or using cribs or electronic aids on exams will result in judicial proceedings in accordance with the University of Hawaii Student Conduct Code. See [http://studentaffairs.manoa.hawaii.edu/policies/conduct\\_code/](http://studentaffairs.manoa.hawaii.edu/policies/conduct_code/) for details.

## Grading and Student Evaluation

- Online homework (ALEKS) will be mandatory and count for a **20% of the final grade**. Instructions on the online homework and the grading system will be given at the beginning of the course.

- Four exams each worth 20% of the final grade. The 4th exam will be held during final exam week and will be 50% cumulative and 50% on the material covered since Exam 3.

Everything will be graded from 0 to 100. *No curves will be applied*. Final grade ranges will be:

<40: F; 40-44: D-; 45-49: D; 50-54: D+; 55-59: C-; 60-64: C; 65-69: C+;  
70-74: B-; 75-79: B; 80-84: B+; 85-89: A-; 90-94: A; 95-100: A+

## Student Disabilities

The University of Hawaii is an equal opportunity/affirmative action institution, dedicated to teaching all students and reaching all learners. It is our commitment to make our lectures and classrooms accessible to all students. If you have a disability and have not voluntarily disclosed its nature and the support you need, you are invited to contact the KOKUA Program of UH (<http://www.hawaii.edu/kokua/>), or talk with the instructor in order to get any accommodation you might need to take the course. This information will be kept confidential. Please do this as early in the course as possible.

## TENTATIVE LECTURE SCHEDULE

1. Ch. 12 Properties of Liquid Solutions
2. Ch. 13 Chemical Kinetics

### **EXAM 1 Friday February 10th**

3. Ch. 14 Chemical Equilibrium
4. Ch. 15 Acids and Bases

### **EXAM 2 Friday March 9th**

5. Ch. 16 Aqueous Ionic Equilibrium
6. Ch. 17 Free Energy and Thermodynamics (w/ some Ch. 6 review)

### **EXAM 3 Friday April 13th**

7. Ch. 18 Electrochemistry

### **FINAL EXAM Friday May 11<sup>th</sup>, 14:15-16:15**