

CHEM 162, Section 2, Spring 2008
GENERAL CHEMISTRY II

Instructor: Prof. Oscar Navarro

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Office hours: Students are encouraged to meet with the instructor for questions, additional information or any other related matter. Office hours are every lecture day, one hour before the lecture and one hour after, no appointment required. Any other time can be scheduled by appointment (request by email). In addition, Bilger 337 will be staffed by Teaching Assistants Monday through Friday. Check the posted schedule for the times when TA's are available.

Classroom: Bilger 152

Hours: 10:30-11:45 Tuesday and Thursday, 3 credit hours

Textbook: "Chemistry, Matter and Its Changes", Brady and Senese, 4th Edition
ISBN: 0-471-44891-5
Study Guide ISBN: 0-471-69247-6

Student Responsibility: It is the student's responsibility to put forth the effort required to learn the material and to become competent with it. The best way of learning is to self-test what you don't know and correct that: this means working on the course, reviewing in a daily basis and using good study habits. Ignoring the subject until three days before the test usually guarantees a failure. If you find yourself having troubles catching up with the subject, the instructor will be happy to help you. Consult him as soon as possible.

Reading the corresponding chapter and trying to understand the concepts before going to the classroom is highly encouraged. Lectures will be presented in PowerPoint, unless otherwise noted. Handouts will be uploaded at the MyUH webpage of the course in order to make them available for the students, so they can be printed out and brought to class to take notes. The handout for the following chapter should be brought in case we advance to it earlier than expected.

Course Policies:

- 1- There will be no makeup exams. If you miss an exam and have a valid excuse, the weighing of the other assignments will be adjusted accordingly.
- 2- Regular attendance in lecture is highly recommended. The aim of the lecture session is to guide you in your studies and to clarify, emphasize and illustrate the important (and sometimes subtle) concepts. Topics not included in the text will be covered in class and will appear in the tests. You are responsible for all information relayed in class whether you attend or not.
- 3- Academic dishonesty will not be tolerated. Cheating in the form of copying, plagiarism, altering information or using cribs on exams will result in judicial proceedings in accordance with the University of Hawaii's policy on academic dishonesty.

Grading and Student Evaluation

Four tests: 20% of the final grade each

Questions (weekly assigned from WileyPLUS): 20% of the total grade

Everything will be graded from 0 to 100. No curves will be applied. Final grades will be:

0-24: F; 25-49: D; 50-55: C-; 56-61: C; 62-66: C+; 67-72: B-; 73-78: B; 79-84: B+; 85-89: A-; 90-95: A; 96-100: A+

Exam results will be available in the tutorial room, Bilger 337, after each examination is graded.

It is mandatory to have a score of **at least 70%** in the assigned WileyPLUS questions.

Student Disabilities

The University of Hawaii is an equal opportunity/affirmative action institution, dedicated to teaching all students and reaching all learners. It is our commitment to make our lectures and classrooms accessible to all students. If you have a disability and have not voluntarily disclosed its nature and the support you need, you are invited to contact the KOKUA Program of UH (<http://www.hawaii.edu/kokua/>, phone (808) 956-7511), or talk with the instructor in order to get any accommodation you might need to take the course. This information will be kept confidential. Please do this as early in the course as possible.

TENTATIVE LECTURE SCHEDULE

TOPICS

1. Ch. 12 Intermolecular Forces, Liquids, Phase Changes
2. Ch. 13 Solids and Their Structure/Properties
3. Ch. 14 Properties of Liquid Solutions

EXAM 1 Thursday February 7th

4. Ch. 20 Thermodynamics (defer secs. 20.8 and 20.9 to topic 5, review Ch.7, secs 7.5-7.8)
5. Ch. 16 Chemical Equilibrium (also read secs. 20.8 and 20.9)

EXAM 2 Thursday February 28th

6. Ch. 17 Acids and Bases (review in Ch. 5, secs. 5.5-5.7)
7. Ch. 18 Acid/Base Equilibria
8. Ch. 19 Solubility Equilibria

EXAM 3 Thursday April 17th

9. Ch. 15 Chemical Kinetics
10. Ch. 21 Electrochemistry

FINAL EXAM Tuesday May 13th, 9:45-11:45